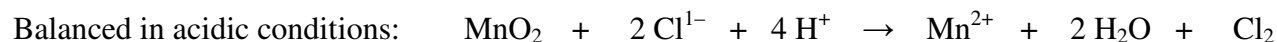
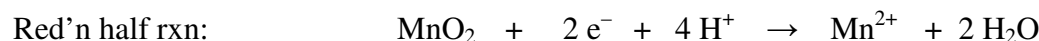
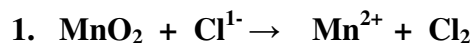
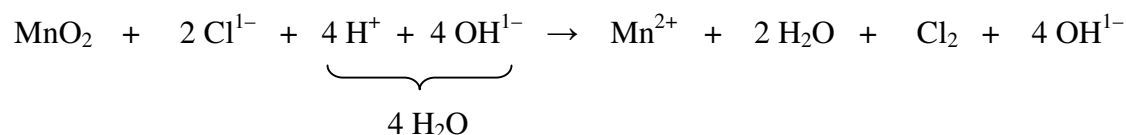


Unit 8, Lesson 04: Answers to Homework

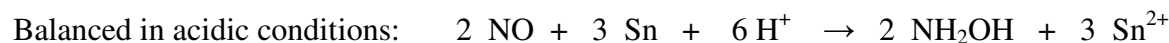
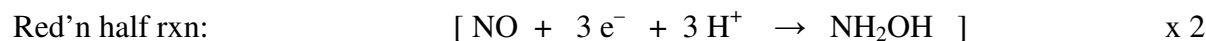
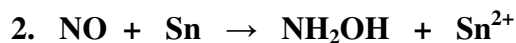
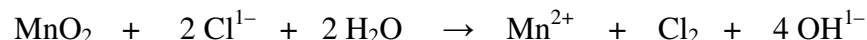
Balancing Redox Reactions in Both Acidic and Basic Conditions from Unit 8 Outline



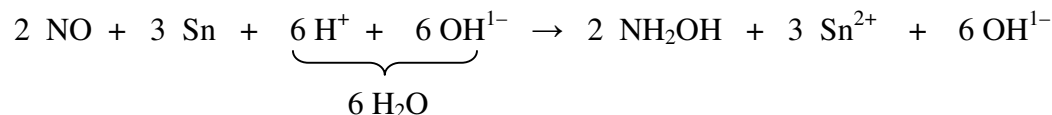
To balance in basic conditions, add 4 OH⁻ ions to both sides:



This will neutralize the acid, and also produce additional water on the left side, so rebalance the water. The reaction balanced in basic conditions is:

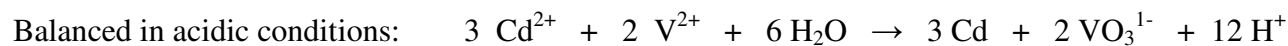
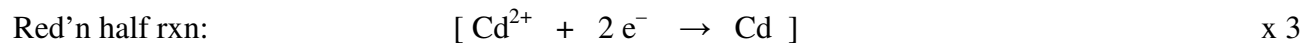
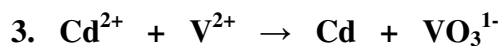


To balance in basic conditions, add 6 OH⁻ ions to both sides:

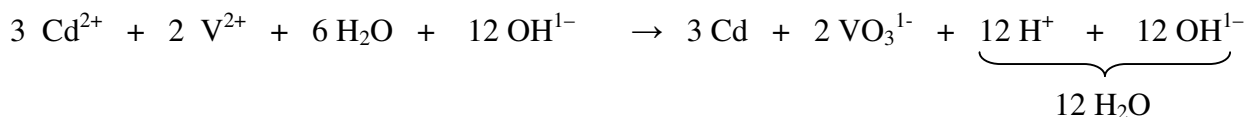


This will neutralize the acid, and also produce water on the left side, giving the reaction balanced in basic conditions:

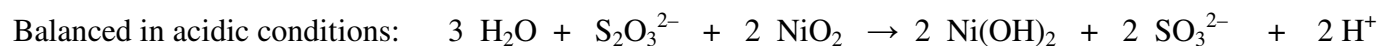
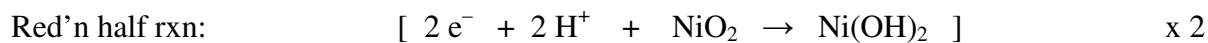
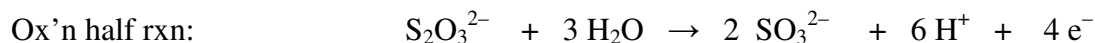
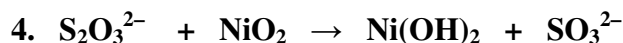




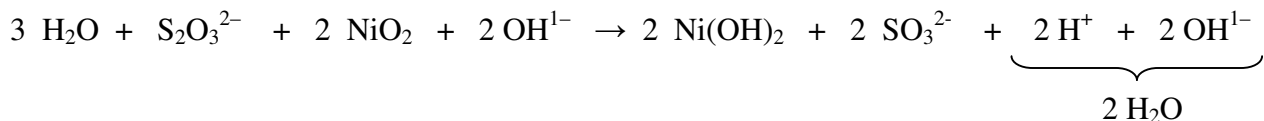
To balance in basic conditions, add 12 OH⁻ ions to both sides:



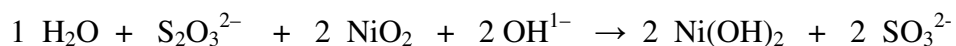
This will neutralize the acid, and also produce water on the right side, giving the reaction balanced in basic conditions:

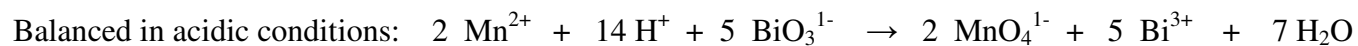
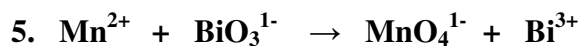


To balance in basic conditions, add 2 OH⁻ ions to both sides:

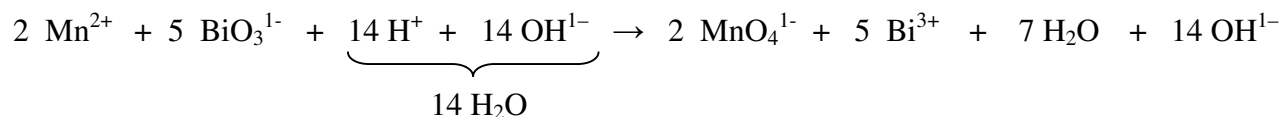


This will neutralize the acid, and also produce water on the right side, giving the reaction balanced in basic conditions:

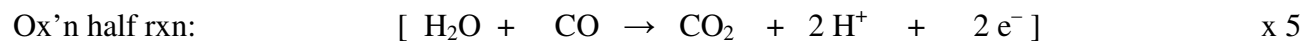
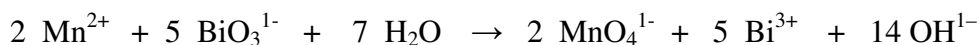




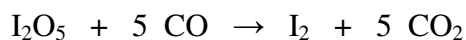
To balance in basic conditions, add 14 OH⁻ ions to both sides:

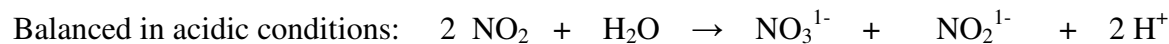
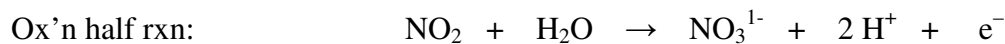


This will neutralize the acid, and also produce water on the left side, giving the reaction balanced in basic conditions:

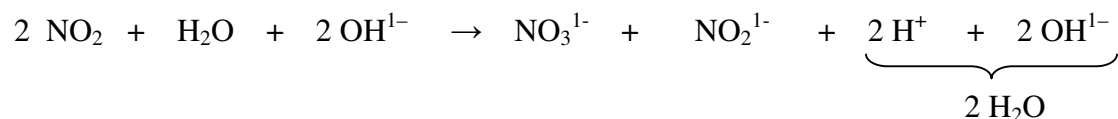


When balanced by adding together, the conditions become neutral. So the reaction can only be balanced in neutral conditions:

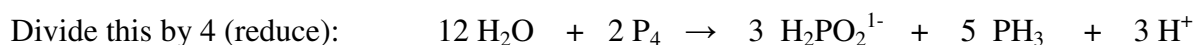
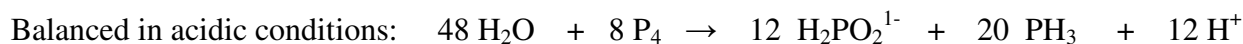
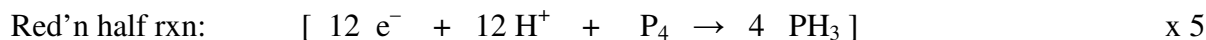




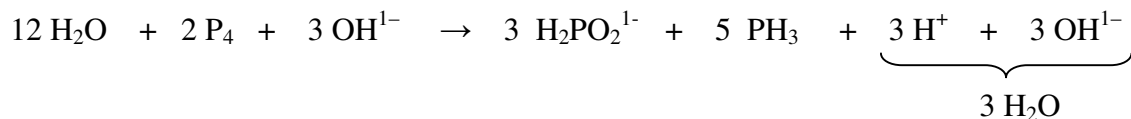
To balance in basic conditions, add 2 OH⁻ ions to both sides:



This will neutralize the acid, and also produce water on the right side, giving the reaction balanced in basic conditions:



To balance in basic conditions, add 3 OH⁻ ions to both sides:



This will neutralize the acid, and also produce water on the right side, giving the reaction balanced in basic conditions:

